

Name: _____

Class: _____

Date: _____

For each chemical formula below, list the number of atoms by type. The first is done for you.

Practice Reading Chemical Formulas

<p>1. H_2O</p> <p>H: <u>2</u></p> <p>O: <u>1</u></p>	<p>2. $NaCl$</p> <p>Na: _____</p> <p>Cl: _____</p>	<p>3. $C_6H_{12}O_6$</p> <p>C: _____</p> <p>H: _____</p> <p>O: _____</p>	<p>4. $PbSO_4$</p> <p>Pb: _____</p> <p>S: _____</p> <p>O: _____</p>	<p>5. Na_2CO_3</p> <p>Na: _____</p> <p>C: _____</p> <p>O: _____</p>
<p>6. $C_7H_5(NO_2)_3$</p> <p>C: _____</p> <p>H: _____</p> <p>N: _____</p> <p>O: _____</p>	<p>7. $Ca(H_2PO_4)_2$</p> <p>Ca: _____</p> <p>H: _____</p> <p>P: _____</p> <p>O: _____</p>	<p>8. $Mg(OH)_2$</p> <p>Mg: _____</p> <p>O: _____</p> <p>H: _____</p>	<p>9. $(NH_4)_2CO_3$</p> <p>C: _____</p> <p>H: _____</p> <p>N: _____</p> <p>O: _____</p>	<p>10. NH_4NO_3</p> <p>N: _____</p> <p>H: _____</p> <p>O: _____</p>
<p>11. $5H_2O$</p> <p>H: _____</p> <p>O: _____</p>	<p>12. $2CaCl_2$</p> <p>Ca: _____</p> <p>Cl: _____</p>	<p>13. $2Na_3PO_4$</p> <p>Na: _____</p> <p>P: _____</p> <p>O: _____</p>	<p>14. $2Fe_2(SO_4)_3$</p> <p>Fe: _____</p> <p>S: _____</p> <p>O: _____</p>	<p>15. $3Mg(NO_3)_2$</p> <p>Mg: _____</p> <p>N: _____</p> <p>O: _____</p>
<p>16. $Ag_2Cr_2O_7$</p> <p>Ag: _____</p> <p>Cr: _____</p> <p>O: _____</p>	<p>17. $3Al_2Si_2O_5(OH)_4$</p> <p>Al: _____</p> <p>Si: _____</p> <p>O: _____</p> <p>H: _____</p>	<p>18. $2Ba(NO_3)_2$</p> <p>Ba: _____</p> <p>N: _____</p> <p>O: _____</p>	<p>19. $2C_4H_{11}NO_3$</p> <p>C: _____</p> <p>H: _____</p> <p>N: _____</p> <p>O: _____</p>	<p>20. $2(NH_4)_3N$</p> <p>N: _____</p> <p>H: _____</p>